

18. Subscript- number placed to the lower right of a chemical symbol to indicate the number of atoms of the element in the compound.
19. Chemical Equation- expression in which symbols, formulas, and numbers are used to represent a chemical reaction.
20. Coefficient- number that is placed in front of a symbol or a formula in a chemical equation that indicates how many atoms or molecules of this substance are involved in the reaction.
21. Emulsion- colloid where one liquid is suspended in another and remains so because of the emulsifier.
22. Emulsifier- substance that keeps on liquid suspended in another. ex. egg yolk is the emulsifier in mayonnaise.
23. Suspension- heterogeneous mixture that settles or separates when allowed to sit undisturbed.
24. Metal- element that is a good conductor of heat and electricity, has a shiny metallic luster, is ductile and malleable. tends to lose electrons when it chemically reacts.
25. Luster- the way a substance reflects light (ex. metallic luster reflects light like a metal, waxy luster reflects light like wax).
26. Ductile- able to be drawn into a thin wire.
27. Malleable- able to be hammered into a thin sheet or bent and retain the altered shape.
28. Nonmetal- an element that is a poor conductor of heat and electricity, has a dull luster is nonmalleable non ductile and tends to gain electrons when it chemically reacts.
29. Metalloid- an element that has properties of both metals and nonmetals.
30. Concentration- measure of amount of solute dissolved in solvent (g/ml).

## Science Chap. 3 Vocabulary

1. Mixture- matter that consists of two or more substances mixed but not chemically combined
2. Heterogeneous Mixture- substance that does not appear to be the same throughout
3. Homogeneous Mixture- substance that appears to be the same throughout
4. Colloid- homogeneous mixture in which the particles are mixed but not dissolved
5. Solution- homogeneous mixture in which one substance is dissolved in the other
6. Solute- substance that is being dissolved in a solution.
7. Solvent- substance that does the dissolving in a solution.
8. Soluble- can be dissolved in another material.
9. Insoluble- can not be dissolved in another material.
10. Solubility- measure of how much solute can be dissolved in a given amount of solvent at a given temperature and pressure.
11. Alloy- a solution of two metals or a metal and a nonmetal that has the properties of a metal.
12. Pure Substance- substance made of only one kind of material, having definite properties, and is the same throughout.
13. Element- simplest type of pure substance, cannot be broken down by light heat or electricity.
14. Atom- the smallest part of an element that has all the properties of that element.
15. Compound- substance made up of molecules that contain more than one kind of atom; two or more elements chemically combined.
16. Molecule- smallest part of a compound that has all the properties of that compound.
17. Chemical Formula- combination of chemical symbols usually used to represent a compound.